

## Intermediate Inorganic Chemistry (CHMC31Y3)

Welcome to CHMC31Y3 course, a course that brings to you the exciting, rich and colorful world of transition elements. Below you will find a more detailed course scope and outline which will, I hope, give you a bit more information about what is in front of us and what is expected from us all this semester.

### Course Scope and Goals

Intermediate Inorganic Chemistry (CHMC31Y3) builds up on the material covered in Introduction to Inorganic Chemistry (CHMB31H3) and will cover topics from the general and special chemistry of transition elements. General topics will include: overview of transition metal properties (their position in the Periodic Table of Elements, relationships to the main group elements, etc.), main classes of compounds, coordination compounds (structure and bonding, general reactivity, magnetic properties), spectroscopic methods in inorganic chemistry (UV, IR, NMR), and basic organometallic chemistry of transition elements. Special topics will include reactivity of some transition metal complexes (through important examples), and catalysis. In general, the course is oriented more towards the physical rather than descriptive inorganic chemistry.

### Course Outline

These are some of the topics that will be covered in the course. We will not cover them necessarily in this order.

#### 1. GENERAL INTRODUCTION TO D- AND F-BLOCK ELEMENTS:

- a. Position in the Periodic Table of Elements, relationship to main group (s- and p- block) elements, electronic configurations;
- b. Physical and chemical properties of the d- and f-block elements

#### 2. COORDINATION COMPOUNDS (OR COMPLEXES):

- a. Definition and brief history;
- b. Ligand Classes;
- c. Coordination numbers and geometries;
- d. Isomers;

#### 3. LIGANDS, STABILITY AND SYMMETRY

- a. Relationships between ligand structure and complex geometry
- b. Complex stability:
  - i. Thermodynamic stability of complexes
  - ii. Chelating and macrocyclic effects

- c. Introduction to symmetry:
    - i. The concept of symmetry
    - ii. Symmetry elements and symmetry operations
    - iii. Point groups
4. BONDING IN COORDINATION COMPOUNDS:
- a. Ligand Field Theory;
  - b. Crystal Field Theory;
  - c. Molecular Orbital Approach;
  - d. Isolobal Analogy.
5. CHARACTERIZATION OF COORDINATION COMPOUNDS I: UV-VIS SPECTROSCOPY
- a. Color of transition metal complexes;
  - b. Spectral terms and selection rules;
  - c. Correlation diagrams;
  - d. Charge transfer: metal-to-ligand and ligand-to-metal.
6. CHARACTERIZATION OF COORDINATION COMPOUNDS II: INFRARED SPECTROSCOPY (THEORY AND APPLICATIONS);
7. CHARACTERIZATION OF COORDINATION COMPOUNDS III: NMR SPECTROSCOPY:
- a. General introduction to NMR spectroscopy;
  - b. NMR active nuclei;
  - c. Chemical shifts, coupling constants and fundamentals of interpretation of NMR spectra of coordination and organometallic compounds (NMR in inorganic vs. organic chemistry);
  - d. Fluxional compounds.
8. REACTIVITY OF COORDINATION COMPOUNDS:
- a. General introduction (thermodynamics, kinetics, mechanisms);
  - b. Substitution reactions:
    - i. In square planar complexes,
    - ii. In octahedral complexes;
  - c. Isomerization;
  - d. Electron-transfer mechanisms.

## 9. ORGANOMETALLIC COMPOUNDS:

- a. Ligands in organometallic chemistry;
- b. 18-electron rule and structure of organometallic compounds.
- c. Basic classes of organometallic compounds:
  - i.  $\sigma$ -bonded alkyl and aryl complexes
  - ii.  $\pi$ -bonded systems (alkenes, alkynes, cyclopentadienyl and other aromatic systems)
  - iii. Other common ligands in organometallic chemistry: hydride, dihydrogen, and phosphines

## 10. SPECIAL TOPICS I: CATALYSIS - CHEMISTRY AND INDUSTRY:

- a. Energy considerations, green chemistry and atom economy principles;
- b. Heterogeneous catalysis;
  - i. Principles;
  - ii. Mechanisms;
  - iii. Examples;
- c. Homogeneous catalysis;
  - i. Principles;
  - ii. Mechanisms;
  - iii. Examples;
- d. Homogeneous vs. heterogeneous catalysis: which way to go?
- e. Industry.

## 11. SPECIAL TOPICS II: BIOINORGANIC COORDINATION CHEMISTRY

- a. The elements of life: s-, p- and d-block elements in living systems
- b. Criteria for element selection: abundance, availability and usefulness
- c. Most important biological ligands: amino acids, corrins and small inorganic molecules d. Metal – protein symbiosis in living systems

### Evaluation:

Literature assignments	10%
Midterm Exam	20%
Final Exam	35%
Lab component	35%

### Class participation:

If you attend lectures and participate in class discussions by answering or asking questions

- Your mark will be rounded to the next GPA level (e.g. 80 %, 70 %, etc.) **IF AND ONLY IF** your final mark is just one percentage point below the next GPA cut-off (e.g. 79 %, 69 %, etc.)
- If you do not attend lectures and do not participate in class discussions, then this bonus will not apply.

The dates for the exams and literature assignment will be announced during labs or classes. What follows are some details regarding the content of each.

*Literature assignments.* The “assignments” are based on the experiments. The detailed description of the assignment parts is provided in the lab manual introduction.

*Midterm Exam.* The date/location of the midterm exam will be announced once the campus-wide scheduling of midterms has been completed. The duration of this exam is 90 minutes (1.5 h). The exact material on the midterm exam will depend on the day the exam is scheduled. Any aids that you might need (i.e. calculators and such) will be announced in advance of the exam. As mentioned above, the knowledge of material from both the lecture notes and your textbook is expected and as such is testable material.

*Final exam.* The final exam is cumulative meaning that material covered before and after midterm is going to be on the exam. The duration of the final exam is 180 minutes (3 hours).

Past exam questions will be provided and posted on the blackboard, but *no answers* will be provided. If you need help – come and ask!

*Lab component.* For more information on lab component see “Introduction” part of the lab manual.

Since you have to submit your formal lab report to turnitin.com, please make a note of the following:

Normally, students will be required to submit their course essays to Turnitin.com for a review of textual similarity and detection of possible plagiarism. In doing so, students will allow their essays to be included as source documents in the Turnitin.com reference database, where they will be used solely for the purpose of detecting plagiarism. The terms that apply to the University's use of the Turnitin.com service are described on the Turnitin.com web site.

#### *Missed term test policies*

If you **miss a term test**, you must provide the appropriate documentation **within one week of the term test**. If the reason is medical, you should download the official UTSC medical form available at [http://www.utsc.utoronto.ca/~registrar/resources/pdf\\_general/UTSCmedicalcertificate.pdf](http://www.utsc.utoronto.ca/~registrar/resources/pdf_general/UTSCmedicalcertificate.pdf) and have your doctor complete the form. If no acceptable documentation for your absence is provided (again: within one week of the test) you will be assigned zero grade for that test. With the documentation the value of the missed test will be added to your final exam (for example, if you miss a term test for a valid reason and chose this option, your final exam will be worth  $35\% + 20\% = 55\%$  of your final mark).

#### **Office hours and contact info**

My office is located in the environmental sciences and chemistry building (behind the Instructional Center), 5<sup>th</sup> floor, room EV546. **The office hours schedule will be posted on the Blackboard portal (under ‘Contact’)**. If you would like to see me outside the office hours (for any reason), please e-mail me and we’ll schedule the time. You can pay me a visit before the semester starts and before announcement of the regular office hours.

I can also be reached via e-mail: m.zimmer.deiuliis@utoronto.ca.

## Suggested reading materials

Your lecture notes, which will be available on the Blackboard, should be your major guides to mastering the material for this course. **However, the knowledge of both textbook and lecture materials is required.** Apart from the required textbook below, we shall analyze selected journal articles relevant to the course materials. The required references will be posted on the blackboard as citations.

*Our required textbook is the same as for CHMB31H3:*

Atkins, Overtone, Rouke, Weller, Armstrong and Hagerman. **Shriver and Atkins' Inorganic chemistry.** 6<sup>th</sup> edition. New York: W.H. Freeman and Company, 2014.

These are additional sources for those of you who would like to explore more and can be found in the library (some texts are available in electronic format through the UofT library catalogue):

Wilkinson, A. and A. Cotton. Advanced Inorganic Chemistry. 5<sup>th</sup> ed. New York; Toronto: Wiley, 1988 (a very detailed descriptive inorganic chemistry – for those who need or would like to learn more about the chemistry of elements)

Greenwood, N.N. and A. Earnshaw. Chemistry of the Elements. 2<sup>nd</sup> ed. Oxford: Butterworth Heinemann, 1998. (Probably one of *the best* and most detailed descriptive inorganic chemistry textbook out there, but does not cover in great detail spectroscopic techniques and bonding. Useful to learn a lot about the elements and their reactivity).

Miessler, G.L., and D.A. Tarr. Inorganic Chemistry. 3<sup>rd</sup> ed. Upper Saddle River: Pearson Prentice Hall, 2004. (A good text for our topics 2, 3 (UV-Vis), 6, and 8; it is on course reserves in UTSC library)

Huheey, J.E., E.A. Keiter, and R.L. Keiter. Inorganic Chemistry: Principles of structure and reactivity. 4<sup>th</sup> ed. Upper Saddle River: Pearson Prentice Hall, 1993-94 (a classic textbook, covers many relevant topics for our course)

Crabtree, R. H. The Organometallic Chemistry of Transition Metals. 4<sup>th</sup> ed. Wiley-Interscience, 2005 (Useful the organometallic topics)

The additional readings are not required materials, but are sources that can provide you with more detail on the topics you would like to explore further on your own.

## Academic Integrity

Academic integrity is one of the cornerstones of the University of Toronto. It is critically important both to maintain our community which honors the values of honesty, trust, respect, fairness and responsibility. It also protects you, the student within our community as well as the value of the degree towards which you are all working so diligently. Detailed information about how to act with academic integrity, the Code of Behavior on

Academic Matters, and the processes by which allegations of academic misconduct are resolved can be found online: <http://www.artsci.utoronto.ca/osai/students> and [http://www.utsc.utoronto.ca/~vpdean/academic\\_integrity.html](http://www.utsc.utoronto.ca/~vpdean/academic_integrity.html)

Section B of the University of Toronto's Code of Behaviour on Academic Matters (<http://www.governingcouncil.utoronto.ca/policies/behaveac.htm>) lists actions that are considered academic offences. Some of the most common offences are:

- To use someone else's ideas or words in their own work without acknowledging that those ideas/words are not their own with a citation and quotation marks, i.e. to commit **plagiarism**.
- To include **false, misleading** or **concocted** citations in their work.
- To obtain **unauthorized** assistance on any assignment.
- To provide **unauthorized** assistance to another student. *This includes showing another student completed work.*
- To submit their own work for credit in **more than one course** without the permission of the instructor.
- To **falsify** or **alter** any documentation required by the University. This includes, but is not limited to, doctor's notes.
- To use or possess an unauthorized aid in any test or exam.

There are other offences covered under the Code, but these are by far the most common. Please respect these rules and the values which they protect. Offences against academic integrity will be dealt with according to the procedures outlined in the Code of Behavior on Academic Matters.

## **Accessibility**

Students with diverse learning styles and needs are welcome in this course. In particular, if you have a disability/health consideration that may require accommodations, please feel free to approach me and/or the AccessAbility Services Office ([ability@utsc.utoronto.ca](mailto:ability@utsc.utoronto.ca)) as soon as possible. I will work with you and AccessAbility Services to ensure you can achieve your learning goals in this course. Enquiries are confidential. The UTSC AccessAbility Services staff are available by appointment to assess specific needs, provide referrals and arrange appropriate accommodations. More details are available at: <http://www.utsc.utoronto.ca/~ability/>.